Sander, R., 1999, Compilation of Henry's Law Constants for Inorganic and Organic Species of Potential Importance in Environmental Chemistry, version 3, 8 Apr 1999
http:\\www.mpch-mainz.mpg.de\~sander\res\henry.html

Errata (collected from the above web site 4 June 2003)

Well, I knew it would happen... A big list like this just couldn't be perfect.
I have now added a list of errors that I have made in my compilation.

Errata for version 3 (17 Feb 99)

In note 20, the solubility for strong acids should be defined as:

$kH = ([H+] * [A-]) / p(HA)$

and NOT

$kH = ([H+] + [A-]) / p(HA)$

Thanks to J. C. Wheeler for pointing this out to me.

There's a nasty little error in the sign of the temperature dependence in all tables:
The values given in the third column are

$d \ln kH / d (1/T)$

and NOT

$- d \ln kH / d (1/T)$

The solubility decreases with increasing temperature, of course. This is shown correctly in equation (5) on page 3.

Data from Allen et al. (1998) are measured, not estimated.

The reference Gmehling et al. (1981) should be: J. Gmehling and U. Onken and W. Arlt. Vapor-Liquid Equilibrium Data Collection. D. Behrens and R. Eckermann (eds.). Dechema, Frankfurt/Main Vol. 1a (1981)